

FIFTH SEMESTER B.Sc. DEGREE EXAMINATION, NOVEMBER 2019

(CUCBCSS-UG)

Chemistry

CHE 5B 08—PHYSICAL CHEMISTRY—II

Time : Three Hours

Maximum : 80 Marks

Section A

*Answer all questions.**Each question carries 1 mark.*

1. Example of a molecule belonging to C_{3v} point group is _____.
2. The number of components and variance of the following system is _____.

Ice \rightleftharpoons water \rightleftharpoons water vapour.

3. The number of vibrational modes possible for SO_2 is _____.
4. Which among the carbon isotopes has a nuclear spin ?

$${}^6_{11}C, {}^6_{12}C, {}^6_{13}C \text{ and } {}^6_{14}C.$$

5. Give an example of Photochemical reaction.
6. Calculate the energy of an Einstein of radiation of wavelength 250 nm.
7. R_f value is defined as _____.
8. Catalyst used in Zeigler-Natta polymerisation is _____.
9. Arrhenius equation which expresses the variation of rate constant of a reaction with temperature is _____.
10. An example of a molecule with centre of inversion is _____.

(10 × 1 = 10 marks)

Section B

*Answer any ten questions.**Each question carries 2 marks.*

11. State mutual exclusion rule. Illustrate with example.
12. Write a note on azeotropic mixtures.
13. How does temperature influence the rate of a reaction ? Explain.
14. State Stark. Einstein law and explain the term Quantum yield of a photochemical reaction.
15. Write S.N. on Thin layer chromatography.

Turn over

16. What is Dorn effect ?
17. Briefly discuss chemisorption.
18. Define proper axis and improper axis of symmetry.
19. State and explain Frank–Condon principle.
20. What is meant by (i) Finger print region ; (ii) Chemical shift ?
21. Distinguish between adsorption and absorption.
22. Write S.N. as phosphorescence.

(10 × 2 = 20 marks)

Section C

*Answer any five questions.
Each question carries 6 marks.*

23. Write S.N. on (a) Electrical double layer ; (b) Protective colloids.
24. Define group and point group. construct GMT for C_{2v} point group.
25. Discuss theory of Homogenous and Heterogeneous catalysis.
26. Describe the collision theory of reaction rates.
27. Explain Pattinson's process of desilverization of lead.
28. State and explain Nernst distribution law.
29. Discuss principle, process and applications of Gas chromatography.
30. What are the rules that members of a group must obey ??

(5 × 6 = 30 marks)

Section D

*Answer any two questions.
Each question carries 10 marks.*

31. (a) Write S.N. as Jablonski diagram.
(b) Optical and electrical properties of colloids.
32. (a) Write S.N. on (i) photochemical Hydrogen–Bromine reaction ; (ii) BET equation.
(b) Draw the phase diagram of water system and discuss the application of phase rule to the system.
33. (a) Briefly discuss Langmuir and Freundlich Isotherms.
(b) Upper CST and lower CST.
34. (a) Distinguish between Emulsions and gels.
(b) Write S.N. as NMR spectroscopy.
(c) Stoke's and antistoke's lines.

(4 + 4 + 2 = 10 marks)

[2 × 10 = 20 marks]